

PART 1 - CHEMISTRY

Topic : Electrochemistry

SECTION-B : (Maximum Marks : 4)

This section contains **ONE** question.

The answer to each question is a **NON-NEGATIVE INTEGER**.

For each question, enter the correct integer corresponding to the answer in the place designated to enter the answer.

Answer to each question will be evaluated according to the following marking scheme :

Full Marks : +4 If **ONLY** the correct answer is entered;

Zero Marks : 0 In all other cases.

1. In an electrochemical cell, dichromate ions in aqueous acidic medium are reduced to Cr^{3+} . The current (in amperes) that flows through the cell for 48.25 minutes to produce 1 mole of Cr^{3+} is _____.

Use: 1 Faraday = 96500 C mol^{-1}

[JEE Advanced 2025 Paper 1]

Scan or tap for the
full solution



Space for Rough Work

JEE ADVANCED ELECTROCHEMISTRY — FARADAY'S LAWS OF ELECTROLYSIS — PYQS

Test Pattern : **JEE (Advanced)**

ANSWER KEY

[illegible]