

PART 1 - CHEMISTRY

Topic : Equilibrium

SECTION-B : (Maximum Marks : 4)

This section contains **ONE** question.

The answer to each question is a **NUMERICAL VALUE**.

For each question, enter the correct numerical value of the answer (in decimal notation, truncated / rounded-off to the **SECOND DECIMAL PLACE**; e.g. 6.25, 7.00, -0.33, -.30, 30.27, -127.30) in the place designated to enter the answer.

Answer to each question will be evaluated according to the following marking scheme :

Full Marks : +4 If **ONLY** the correct answer is entered;

Zero Marks : 0 In all other cases.

1. The solubility of barium iodate in an aqueous solution prepared by mixing 200 mL of 0.010 M barium nitrate with 100 mL of 0.10 M sodium iodate is $X \times 10^{-6}$ mol dm⁻³. The value of X is _____.

Use: Solubility product constant (K_{sp}) of barium iodate = 1.58×10^{-9}

[JEE Advanced 2025 Paper 2]

Scan or tap for the
full solution



Space for Rough Work

**JEE ADVANCED EQUILIBRIUM — SOLUBILITY
PRODUCT (KSP) AND PRECIPITATION — PYQS**

Test Pattern : JEE (Advanced)

ANSWER KEY**SECTION-B**

Ques.	1									
Ans.	3.85 to 4.15									