

PART 1 - CHEMISTRY

Topic : Equilibrium

SECTION-B : (Maximum Marks : 4)

This section contains **ONE** question.

The answer to each question is a **NUMERICAL VALUE**.

For each question, enter the correct numerical value of the answer (in decimal notation, truncated / rounded-off to the **SECOND DECIMAL PLACE**; e.g. 6.25, 7.00, -0.33, -.30, 30.27, -127.30) in the place designated to enter the answer.

Answer to each question will be evaluated according to the following marking scheme :

Full Marks : +4 If **ONLY** the correct answer is entered;

Zero Marks : 0 In all other cases.

1. At 25 °C, the concentration of H^+ ions in 1.00×10^{-3} M aqueous solution of a weak monobasic acid having acid dissociation constant (K_a) of 4.00×10^{-11} is $X \times 10^{-7}$ M. The value of X is _____.

Use: Ionic product of water (K_w) = 1.00×10^{-14} at 25 °C

[JEE Advanced 2025 Paper 1]

Scan or tap for the
full solution



Space for Rough Work

**JEE ADVANCED EQUILIBRIUM — PH OF VERY WEAK
ACID INCLUDING WATER IONISATION — PYQS**

Test Pattern : JEE (Advanced)

ANSWER KEY**SECTION-B**

Ques.	1												
Ans.	2.2 to 2.3												