

PART 1 - CHEMISTRY

Topic : Solutions

SECTION-B : (Maximum Marks : 8)

This section contains **ONE** paragraph.

Based on each paragraph, there are **TWO** questions.

The answer to each question is a **NUMERICAL VALUE**.

For each question, enter the correct numerical value of the answer (in decimal notation, truncated / rounded-off to the **SECOND DECIMAL PLACE**; e.g. 6.25, 7.00, -0.33, -.30, 30.27, -127.30) in the place designated to enter the answer.

Answer to each question will be evaluated according to the following marking scheme :

Full Marks : +4 If **ONLY** the correct answer is entered;

Zero Marks : 0 In all other cases.

Paragraph for Questions 1 and 2

Two volatile liquids A and B form an ideal solution. Consider a 5 molal solution of B in A inside a closed container having a total vapour pressure of 100 mm Hg at 300 K. The vapour pressure of pure A at 300K is 105 mm Hg. Assume that A and B behave as ideal gases in the vapour phase.

Given:

The gas constant $R = 0.08 \text{ L atm K}^{-1} \text{ mol}^{-1}$

Molar mass of A is 50 g mol^{-1}

Molar mass of B is 57 g mol^{-1}

Density of liquid B at 300 K is 0.5 g/mL

$1 \text{ atm} = 760 \text{ mm Hg}$

- At 300 K, the ratio of the molar volume of pure B in vapour phase to its molar volume in liquid phase is ____.

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Scan or tap for the
full solution 

- The mole fraction of B in vapour phase which is in equilibrium with this solution is ____.

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Scan or tap for the
full solution 

Space for Rough Work

**JEE ADVANCED SOLUTIONS — RAOULT'S LAW,
VAPOUR PRESSURE — PYQS**

Test Pattern : JEE (Advanced)

ANSWER KEY**SECTION-B**

Ques.	1	2							
Ans.	1991 to 2053	0.16							