

PART 1 - CHEMISTRY

Topic : Thermodynamics

SECTION-B : (Maximum Marks : 4)

This section contains **ONE** question.

The answer to each question is a **NUMERICAL VALUE**.

For each question, enter the correct numerical value of the answer (in decimal notation, truncated / rounded-off to the **SECOND DECIMAL PLACE**; e.g. 6.25, 7.00, -0.33, -.30, 30.27, -127.30) in the place designated to enter the answer.

Answer to each question will be evaluated according to the following marking scheme :

Full Marks : +4 If **ONLY** the correct answer is entered;

Zero Marks : 0 In all other cases.

1. Considering ideal gas behavior, the expansion work done (in kJ) when 144 g of water is electrolyzed completely under constant pressure at 300 K is _____.

Use: Universal gas constant (R) = $8.3 \text{ J K}^{-1} \text{ mol}^{-1}$; Atomic mass (in amu): H = 1, O = 16

[JEE Advanced 2025 Paper 1]

Scan or tap for the
full solution



Space for Rough Work

**JEE ADVANCED THERMODYNAMICS — EXPANSION
WORK, $\Delta n_{\text{gas}} RT$ — PYQS**

Test Pattern : JEE (Advanced)

ANSWER KEY**SECTION-B**

Ques.	1								
Ans.	-29.95 to -29.8								