

PYQSide

NEET CHEMICAL KINETICS — INTEGRATED RATE LAW OF ZERO ORDER REACTION — PYQS

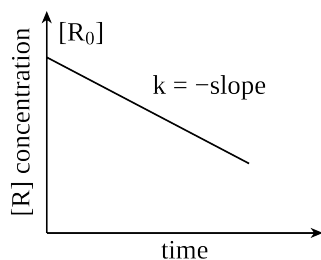
Max. Marks: 12

Questions: 3

Section A (MCQ)

- Q.1** For a certain reaction $R \rightarrow$ Product, the plot of concentration $[R]$ vs time has a negative slope as shown. The order of reaction is :

[NEET (UG) 2026]



- (1) 1 (2) 2.5 (3) 2 (4) 0

Scan or tap for the
full solution



- Q.2** $2A \xrightarrow{k} B$ is a zero-order reaction, where $k = 1.0 \text{ mol L}^{-1} \text{ min}^{-1}$. If the initial concentration of A is 2 M, then the time taken to complete 75% of the reaction will be

[RE-NEET (UG) 2026]

- (1) 2.0 min (2) 1.5 min
(3) 0.75 min (4) 1.0 min

Scan or tap for the
full solution



- Q.3** The rate constant of the reaction $A \rightarrow B$ is $0.6 \times 10^{-3} \text{ mol L}^{-1} \text{ s}^{-1}$. If the concentration of A is 5 M, then concentration of B after 20 minutes is

[AIPMT 2015]

- (1) 3.60 M (2) 0.36 M (3) 0.72 M (4) 1.08 M

Scan or tap for the
full solution



ANSWER KEY

Ques.	1	2	3								
Ans.	4	3	3								